



Iron

Ferrozin

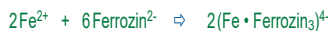
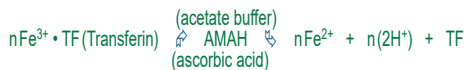
General Information

Characteristics

- Stable at room temperature, without corrosive reagents.
- Without foaming detergents – hence simple and clean pipetting resulting in high precision.
- Automatable due to short reaction time.
- Complete separation of iron from transport protein.
- Without protein removal, thus simple working steps and small sample volume.
- High extinction and good linearity: up to 180 µmol/l Fe (= 1000 µg/dl).
- Evaluation by extinction coefficient or by standard.
- Quality control with conventional control sera (also of animal origin).

Principle

In acetate buffer transferrin (TF) bound iron dissociates into ferric ions which are reduced to ferrous ions (Fe²⁺) in the presence of ascorbic acid. AMAH promotes the reaction. The ferrous ion reacts with the chromogen Ferrozin to form a blue color complex (chelate) with an extinction maximum at 562 nm. The method is specific for iron in serum/plasma and the absorbance is directly proportional to the serum iron concentration up to 180 µmol/L (1000 µg/dL).



Reagents

The reagents are originally closed and stored at room temperature (max. +25 °C) stable until the expiration date.

Risks and Safety

Please observe the necessary precautions for use of laboratory reagents and body fluids. Applications should be performed by expert personnel only. Follow the national and laboratory internal guidelines for work safety and infection control. Wear suitable protective clothing and disposable gloves while handling.

It is important to ensure effective protection against infection according to laboratory guidelines.



For additional safety information please refer to the information on the label and the corresponding Safety Data Sheet (SDS).

Download by QR-Code or link:

- www.sds-id.com/100055-9 (R1a = Iron Buffer)
- www.sds-id.com/100056-8 (R1b = Reducing Reagent)
- www.sds-id.com/100057-7 (R2 = Chromogenic Reagent Ferrozin)
- www.sds-id.com/100059-5 (CAL = Calibrator (Standard))

Main Components

006511-...	Cont.	4.50 mol/L AMAH, 1.55 mol/L acetate, 25 mmol/L thiocarbamide
006512-...	Cont.	Ascorbic acid cryst.
006513-...	Cont.	40 mval/l Ferrozin (buffered pH = 4,5)
006516-...	Cont.	25.0 µmol/L = 140 µg/dl Fe ³⁺ (buffered)

006502-6001	KIT	6x 25 mL Iron Ferrozin
006511-0025	R1a	6x 25 mL Iron Buffer
006512-0025	R1b	6x 50 mg Reducing Reagent
006513-0006	R2	1x 6 mL Ferrozin (Chromogen)
006516-0010	CAL	1x 10 mL Calibrator (Standard)

006502-6002	KIT	4x 100 mL Iron Ferrozin
006511-0100	R1a	4x 100 mL Iron Buffer
006512-0025	R1b	4x 200 mg Reducing Reagent
006513-0016	R2	1x 16 mL Ferrozin (Chromogen)
006516-0010	CAL	1x 10 mL Calibrator (Standard)

Spezimen

Serum, heparin plasma: stable at + 4 °C for 7 days, at room temperature for 4 days.

Do not use hemolytic samples.

Preanalytics

Centrifuge samples immediately. When collecting blood, do not use the first 2 ml for Fe determination, as Fe residues may be aspirated from the disposable cannula. This can lead to increased Fe values!

Only use iron-free disposable material.

Reference Ranges

	[µmol/l]	[µg/dl]
Males:	9.5 ... 29.9	53 ... 167
Females:	8.8 ... 27.0	49 ... 151

Increased iron levels may be caused by hormonal contraceptives. The iron level in serum fluctuates day to day up to 30 %, during the day up to 32 %. In the morning the iron level is higher than in the evening. The iron level decreases with aging.

Diagnostic relevance

Due to the high fluctuations of the iron level in sera from day to day, diagnostic conclusions should be drawn only after repeated testing with similar results. A single analysis has only low diagnostic relevance. A useful diagnostic method to complete the analysis of iron in serum is to test the total and latent (unsaturated) iron-binding capacity (TIBC, UIBC).

Diseases, causing decreased iron level:

- Acute blood loss due to severe external or internal bleeding
- Chronic blood loss (e. g. occult gastrointestinal bleeding)
- Menstrual blood loss
- Hookworm infection (Ankylostoma duodenale)
- Iron absorption disorder (e. g. celiac disease)
- Bacterial infections, inflammations, malign tumors
- Nephrotic syndrome, exudative enteropathy, atransferrinemia
- Malnutrition (Fe-uptake < 10 mg/day), increased need for iron (e. g. due to pregnancy, during growth)

Diseases, causing increased iron level:

- Necrotic liver parenchyma
- Pancreatic insufficiency
- Thalassemia major
- Vitamin B6 deficiency
- Plumbism
- Idiopathic hemochromatosis

Quality control

To control precision and accuracy it is recommended to use a high-quality control serum.

Notes

Centrifuge samples immediately. When taking the blood sample Do not use the first 2 mL for iron analysis, due to possible iron contamination from the disposable needle. This may cause increased results!

Use iron-free disposables only.

Classifications

Not for human diagnostics.

Support/Infoservice

For methodological and technical support, please contact us by E-Mail at support@bioanalytic.de.

Periodically check for updates of this product information on our website.

Feedback

Information from users can be reported to support@bioanalytic.de.

Suggestions for further developments will be considered.

Waste Management

Please observe your national laws and regulations.

Used and expired solutions must be disposed of in accordance with your local regulations.

Inside the EU, national regulations apply that are based on the current, amended version of Council directive 67/548/EEG on the approximation of the laws, regulations and administrative provisions relating to the classification, packaging and labelling of dangerous substances.

Decontaminated packaging can be disposed of as household waste or recycled, unless otherwise specified.

Ordering Information

Calibrator (Fe-Standard) and Chromogenic Reagent Ferene may not be sufficient for adaptation on automatic analyzers and are also available separately if required.

Literature & Footnotes

Legends for the graphic symbols and tags used follow relevant norms or are available on our internet pages.

- [1] L. Thomas, Labor und Diagnose.



Iron

Ferrozin

Sample / Sample Blank - Technique, Manual, Endpoint

Method

Sample/Sample Blank technique, manual, endpoint. Ferrozin/ascorbic acid method without protein removal.

Product information for measuring of iron in serum and plasma. Evaluation by extinction coefficient or by standard. Automatable due to short reaction time.

Attention!

This additional information is a supplement to the product information. It is also important to observe the information in the product information!

Preparation

R1 Buffer/Blank

For preparation of reducing buffer R1 dissolve the content of one tube reducing reagent 6512 (R1b) in one bottle of iron buffer 6511 (R1a).

Shelf life: stable at +2...+8 °C for 14 days, at +15...+25 °C for 5 days

R2 Buffer/Chromogen

Add 1 mL chromogenic reagent 6513 per 25 mL R1 (Reducing buffer) and mix carefully.

Add 4 ml 6513 colour reagent per 100 ml R1 (sample blank reagent) and mix carefully.

Shelf life: see R1.

Procedure

Wavelength:..... 562 nm, Hg 578 nm

Optical path length:..... 10 mm

Temperature:..... +20...+37 °C

measuring mode:..... against sample blank value

Add to reaction tube or cuvette:

	RB	CAL	SB	SA
AQ Aqua p. a.	100 µl	—	—	—
CAL Standard	—	100 µl	—	—
SA Sample	—	—	100 µl	100 µl
R1 Buffer/Blank	—	—	500 µl	—
R2 Buffer/Chromogen	500 µl	500 µl	—	500 µl

Mix and measure after 1...60 minutes sample (E_{SA}) against sample blank (E_{SB}).

The reagent blank value (E_{RB}) must be subtracted from the measured values (determined once per series).

Evaluation/Calculation

$$E_{SA} - E_{SB} - E_{RB} = \Delta E_{SA}$$

1. by extinction coefficient (Hg 578 nm):

$$\mu\text{mol/L Fe} = \Delta E_{SA} \times 238$$

$$\mu\text{g/dL Fe} = \Delta E_{SA} \times 1330$$

2. by extinction coefficient (562 nm):

$$\mu\text{mol/L Fe} = \Delta E_{SA} \times 218$$

$$\mu\text{g/dL Fe} = \Delta E_{SA} \times 1220$$

3. by standard:

$$\mu\text{mol/L Fe} = \Delta E_{SA} \times (25 / E_{CAL})$$

$$\mu\text{g/dL Fe} = \Delta E_{SA} \times (140 / E_{CAL})$$

4. Conversion:

$$\mu\text{g/dL Fe} = \mu\text{mol/L} \times 5.59$$

$$\text{mg/L [ppm] Fe} = \mu\text{mol/L} \times 0.0559$$

Nomenclature

- CAL = Calibrator (Standard)
- E_{SA} = Extinction/Absorbance Sample
- E_{SB} = Extinction/Absorbance Sample Blank
- E_{RB} = Extinction/Absorbance Reagent Blank
- E_{ST} = Extinction/Absorbance Standard



Iron

Ferrozin

E₁ / E₂ - Technique, Manual, Endpoint

Method

E₁/E₂ technique, manual, endpoint. Ferrozin/ascorbic acid method without protein removal.
Product information for manual (single) determination of iron in serum and plasma. Evaluation using extinction coefficients or via standard.
Can be automated due to short reaction time.

Attention!

This additional information is a supplement to the product information. It is also important to observe the information in the product information!

Preparation

- R1 Buffer/Reduction
For preparation of reducing buffer R1 dissolve the content of one tube reducing reagent 6512 (R1b) in one bottle of iron buffer 6511 (R1a).
Shelf life: stable at +2... +8 °C for 14 days, at +15... +25 °C for 5 days
- R2 Ferene (Chromogen)
Solution ready to use.

Procedure

Wavelength:..... 562nm, Hg 578nm
Optical path length:..... 10mm
Temperature:..... +20... +37 °C
Measuring mode:..... E₁/E₂

Add to a cuvette:

SA	Sample	100 µl
R1	Buffer/Reduction	500 µl
R2	Ferrozin (Chromogen)	20 µl

Mix and set photometer to zero (E₁=0). Then add:

Mix and measure after 1...60 minutes (E₂).

The reagent blank value (E_{RB}) must be subtracted from the measured values (determined once per series). For this purpose, Aqua p.a. is used as a sample. For calculation via standard (CAL), this is also treated as a sample.

Evaluation/Calculation

$$E_2 - E_{RB} = \Delta E_{SA}$$

1. by extinction coefficient (Hg 578 nm):

$$\mu\text{mol/L Fe} = \Delta E_{SA} \times 238$$

$$\mu\text{g/dl Fe} = \Delta E_{SA} \times 1330$$

2. by extinction coefficient (562nm):

$$\mu\text{mol/L Fe} = \Delta E_{SA} \times 218$$

$$\mu\text{g/dl Fe} = \Delta E_{SA} \times 1220$$

3. by standard:

$$\mu\text{mol/L Fe} = \Delta E_{SA} \times (25 / E_{CAL})$$

$$\mu\text{g/dl Fe} = \Delta E_{SA} \times (140 / E_{CAL})$$

4. Conversion:

$$\mu\text{g/dl Fe} = \mu\text{mol/L} \times 5.59$$

$$\text{mg/L [ppm] Fe} = \mu\text{mol/L} \times 0.0559$$

Nomenclature

- CAL = Calibrator (Standard)
- E_{SA} = Extinktion/Absorbance Sample
- E_{SB} = Extinktion/Absorbance Sample Blank
- E_{RB} = Extinktion/Absorbance Reagent Blank
- E_{ST} = Extinktion/Absorbance Standard